

CONTRIBUTION FROM THE DEPARTMENTS OF CHEMISTRY OF THE UNIVERSITY OF CALIFORNIA, LOS ANGELES,¹ CALIFORNIA 90024 AND DAVIS, CALIFORNIA 95616

The Crystal and Molecular Structure of 1-Thia-4-telluracyclohexane 4,4-Diiodide, C₄H₈STeI₂

BY CAROLYN KNOBLER,^{2a} J. D. McCULLOUGH,^{2a,b} AND HAKON HOPE^{2c}

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1-Thia-4-telluracyclohexane 4,4-diiodide forms dark purple needles with $a = 18.817$ (12) Å, $b = 9.728$ (8) Å, and $c = 21.746$ (12) Å. The space group is Pbc_a with $Z = 16$ so that there are two molecules of C₄H₈STeI₂ in the asymmetric unit. The intensities of 4296 unique counter-measured reflections were corrected for absorption and used in the structure determination. The phase problem was solved by use of Long's program and the structure was refined by use of full-matrix least-squares procedures to a conventional R index of 5.0% for the 2766 observed reflections. Both molecules have essentially the same structure with approximate mirror symmetry and with the six-membered ring in the chair conformation. The configuration about Te approximates a trigonal bipyramid with I atoms in axial positions and two C atoms plus an unshared electron pair equatorially distributed. The observed bonded distances (in Å) range as follows: Te-I, 2.851 (2) to 2.985 (2); Te-C, 2.13 (2) to 2.18 (2); S-C, 1.79 (3) to 1.88 (3); and C-C, 1.49 (4) to 1.54 (4). Among the bond angles, the more interesting are I-Te-I, 174.9 (1)° and 178.1 (1)°; C-Te-C, 100 (1) and 100 (1)°; C-S-C, 99 (1) and 100 (1)°; and the eight C-Te-I angles, which range from 87.4 (6) to 91.8 (6)°.

Introduction

A study of the structure of 1-thia-4-telluracyclohexane 4,4-diiodide is of interest for two reasons; one, the determination of the structure of the 1,4-thiatellurane ring, and two, a study of the heavy-atom bonding system. An earlier attempt to determine the structure of the parent compound, 1,4-thiatellurane, was complicated by the presence of a type of disorder which gives the molecule a pseudo center of symmetry.³ The project was therefore set aside. Interest in the heavy-atom bonding systems in compounds such as R₂TeI₂ and R₂SeI₂ lies in the varying degrees of combination, through intermolecular bonding, of the axial I-M-I "molecular" type units with the linear M···I-I "complex" type. In both types, bonding systems consisting of nearly linear chains, -I-M-I-I-M-I-, have been found, and within these chains, the M-I and I-I bonded distances vary considerably from one compound to another.⁴⁻⁷ This variability in bond length is probably due to tendencies toward weaker bonding and larger coordination numbers as the atomic numbers of the atoms increase.

Preliminary Study

Crystals of 1,4-thiatellurane 4,4-diiodide were prepared as described by McCullough³ and specimens suitable for the X-ray studies were grown by slow evaporation of solutions in ethylene chloride. The crystals are dark purple needles with a metallic luster and are elongated on the b axis of the orthorhombic unit. The cell constants, based on λ Mo K α_1 0.70926 Å, were found to be $a = 18.817$ (12) Å, $b = 9.728$ (8) Å, and $c = 21.746$ (12) Å at 23°. The systematic extinctions

$0kl$ with k odd, $h0l$ with l odd, and $hk0$ with h odd were noted and these correspond to the space group Pbc_a. The density was observed by flotation to be 3.15 (3) g cm⁻³ which compares favorably with the value 3.134 (5) g cm⁻³ calculated from the X-ray data with $Z = 16$. Thus the crystallographic asymmetric unit consists of two molecules of C₄H₈STeI₂.

The initial intensity data consisted of multiple-film Weissenberg photographs of the levels $h0l$ through $h7l$ and $hk0$ through $hk2$ prepared with Cu K α radiation. The intensities were estimated visually and corrected for absorption ($\mu = 760$ cm⁻¹ for Cu K α) and for the usual geometric factors to give a set of 3765 F_o values. These were used as input to obtain a three-dimensional Patterson summation.⁸ However, interpretation of the summation was not straightforward, so attention was turned to Long's sign prediction program. In this case 222 reflections with $|E| \geq 1.80$ were used. A three-dimensional Fourier summation based on the sign set with the highest consistency index, 0.993, showed all six heavy atoms. A set of structure factors calculated on the basis of two Te and four I atoms with isotropic B values of 3.5 Å² at this point showed an R index of 0.32.

A difference Fourier summation from which the six heavy atom contributions were subtracted showed the two sulfur atoms, but the carbon atoms were not clearly located until another difference map was prepared which was phased on the eight heaviest atoms. The R index at this stage, based on all atoms (except hydrogen) with isotropic B values of 3.5 Å², was 0.20. Two cycles of

(8) Fourier summations were computed by use of UCLAFOUR written by Gantzel and Hope and the least-squares refinement was carried out by use of the full-matrix program ACA No. 317 (Modified) written by Gantzel, Sparks, and Trueblood. This program minimizes $\sum w(\Delta F)^2$ and the weights, w , were assigned by the method of E. W. Hughes, *J. Amer. Chem. Soc.*, **63**, 1737 (1941). The R index was defined as $\sum ||F_o| - |F_c|| / \sum |F_o|$, and the esd values for the parameters were obtained from the final least-squares inverse matrix. The phase problem was solved by use of Long's program, ACA (new) No. 2; thermal ellipsoid plots were prepared by use of Johnson's ORTEP routine and hydrogen positions were calculated by use of HYPSON, written by Hope. The C-H distances were taken as 1.0 Å and the H-C-H angles as 109.5°.

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(2) (a) University of California, Los Angeles; (b) author to whom correspondence should be sent; (c) University of California, Davis.

(3) J. D. McCullough, *Inorg. Chem.*, **4**, 862 (1965).(4) G. Y. Chao and J. D. McCullough, *Acta Crystallogr.*, **14**, 490 (1961).(5) G. Y. Chao and J. D. McCullough, *ibid.*, **15**, 887 (1962).(6) H. Hope and J. D. McCullough, *ibid.*, **17**, 712 (1964).(7) H. Maddox and J. D. McCullough, *Inorg. Chem.*, **5**, 522 (1966).

TABLE I
 ATOMIC POSITIONAL PARAMETERS IN $C_4H_8STeI_2$

(a) Nonhydrogen Atoms			
Molecule A			
	x	y	z
I(1)	0.23678 (6) ^a	0.2655 (2)	0.50188 (6)
I(2)	0.54532 (6)	0.2982 (1)	0.50917 (5)
Te	0.38928 (5)	0.2945 (1)	0.50384 (5)
S	0.4029 (3)	0.1425 (5)	0.3500 (2)
C(1)	0.3972 (11)	0.0790 (17)	0.4766 (10)
C(2)	0.3632 (10)	0.0449 (17)	0.4161 (9)
C(3)	0.3600 (10)	0.3069 (19)	0.3632 (9)
C(4)	0.3893 (11)	0.3883 (19)	0.4156 (9)

Molecule B			
	x	y	z
I(1)	0.09675 (8)	-0.1167 (1)	0.15693 (6)
I(2)	0.10189 (9)	0.4828 (1)	0.16556 (7)
Te	0.10051 (6)	0.1901 (1)	0.15934 (5)
S	0.1875 (3)	0.1543 (7)	0.3083 (2)
C(1)	0.2117 (8)	0.1797 (20)	0.1814 (8)
C(2)	0.2330 (10)	0.2379 (25)	0.2448 (8)
C(3)	0.1023 (11)	0.2357 (24)	0.3019 (8)
C(4)	0.0547 (10)	0.1829 (18)	0.2502 (9)

(b) Calculated Hydrogen Positions

	Molecule A			Molecule B		
	x	y	z	x	y	z
H(1)	0.373	0.022	0.509	0.238	0.236	0.150
H(1)'	0.448	0.054	0.475	0.227	0.083	0.179
H(2)	0.311	0.068	0.418	0.223	0.337	0.246
H(2)'	0.369	-0.055	0.408	0.286	0.221	0.250
H(3)	0.309	0.289	0.372	0.110	0.336	0.294
H(3)'	0.365	0.363	0.325	0.076	0.223	0.341
H(4)	0.361	0.475	0.418	0.010	0.237	0.250
H(4)'	0.440	0.412	0.405	0.043	0.083	0.259

^a Numbers in parentheses here and in succeeding tables are the standard deviations in the least significant digits.

full-matrix least-squares refinement were then computed. The isotropic thermal parameters for I, Te, and S atoms were permitted to refine but the B 's were held fixed at 3.5 \AA^2 for carbon. The R index dropped only slightly, to 0.19. While there was little doubt that the R index could have been further reduced by the introduction of anisotropic thermal parameters in further refinement, an automated diffractometer had become available and we decided to measure a new data set by use of molybdenum radiation. Refinement by use of the copper data was therefore terminated.

Molybdenum Data and Refinement

The crystal selected for the intensity measurements with molybdenum radiation was an elongated (0.3 mm) fragment with a cross section approximating a right triangle with legs of 0.09 and 0.12 mm. It was mounted with the long direction (the crystallographic b axis) along the φ axis of a Picker full-circle diffractometer. The cell dimensions quoted above were obtained from five axis reflections measured at both 2θ and -2θ , and all deviations from mean values were well below 0.010 \AA . The intensities of all 4296 independent reflections with 2θ less than 52° were measured with the diffractometer in automatic mode. The Mo K radiation was filtered through Zr foil and counted with a scintillation detector. The 2θ - θ scan technique (scan speed $2^\circ/\text{min}$) was

 TABLE II
 ANISOTROPIC THERMAL PARAMETERS^a IN \AA^2

Molecule A						
Atom	B_{11}	B_{22}	B_{33}	B_{12}	B_{13}	B_{23}
I(1)	2.54 (4)	6.38 (8)	4.86 (6)	0.62 (5)	-0.25 (4)	-0.29 (5)
I(2)	2.60 (4)	3.45 (5)	3.85 (6)	0.19 (3)	-0.51 (3)	-0.38 (4)
Te	2.42 (4)	2.75 (4)	2.65 (4)	0.25 (3)	-0.50 (3)	-0.38 (3)
S	4.9 (3)	3.8 (2)	3.3 (2)	-0.3 (2)	0.7 (2)	-0.7 (2)
C(1)	4.6 (8)	1.1 (5)	6.2 (9)	-0.5 (6)	-0.9 (8)	-0.6 (6)
C(2)	4.0 (8)	1.9 (6)	4.1 (8)	0.8 (6)	0.2 (6)	0.1 (5)
C(3)	3.8 (8)	3.4 (8)	3.7 (8)	-0.7 (7)	-1.0 (6)	1.1 (6)
C(4)	4.9 (9)	2.6 (7)	4.2 (8)	-0.4 (7)	-1.2 (7)	0.3 (6)

Molecule B						
	B_{11}	B_{22}	B_{33}	B_{12}	B_{13}	B_{23}
I(1)	5.79 (7)	2.65 (5)	4.40 (6)	-0.19 (4)	0.25 (5)	-0.42 (4)
I(2)	6.75 (7)	2.59 (5)	4.88 (6)	-0.13 (5)	-0.64 (6)	-0.07 (4)
Te	3.15 (4)	2.56 (4)	2.52 (4)	-0.03 (4)	-0.56 (3)	-0.08 (3)
S	5.1 (2)	7.3 (3)	3.2 (2)	-2.4 (2)	-1.4 (2)	2.2 (2)
C(1)	1.6 (6)	5.4 (9)	3.4 (7)	-0.3 (6)	-0.6 (5)	-1.2 (7)
C(2)	3.4 (8)	6.9 (9)	3.0 (7)	-1.6 (8)	-0.4 (6)	0.1 (8)
C(3)	5.0 (9)	5.6 (9)	2.4 (7)	-1.9 (8)	-0.2 (6)	0.1 (7)
C(4)	4.1 (8)	3.2 (7)	3.4 (7)	1.3 (7)	1.3 (6)	0.5 (6)

^a The anisotropic temperature factor expression has the form $\exp[-(b_{11}h^2 + b_{22}k^2 + b_{33}l^2 + b_{12}hk + b_{13}hl + b_{23}kl)]$. The relationships between the tabulated B values and the b parameters above are given by $B_{11} = 4a^2b_{11}$, $B_{12} = 2abb_{12}$, etc.

used with the scan range given by $[2\theta(\alpha_1) - 0.8^\circ]$ to $[2\theta(\alpha_2) + 0.8^\circ]$. Background counts (20 sec) were taken at each end of the scan range. The automatic attenuator mechanism was set to keep the count rate below $\sim 10,000$ counts/sec. The intensities of two check reflections were measured about four times a day. These intensities remained constant during the first half of the run, but showed a gradual decline during the second half. The decrease amounted to a maximum of 5% at the end of the run and was ignored in processing the data.

In order approximately to correct the intensity data for absorption ($\mu = 95 \text{ cm}^{-1}$ for Mo K α), the variations in intensity of the $0k0$ reflections as a function of φ were measured. The reflections 040, 060, and 0,10,0 gave the same φ dependence within $\pm 1\%$ while 020 showed a somewhat larger variation. Absorption corrections were applied to all reflections by use of the mean curve for 040, 060, and 0,10,0, with the maximum correction amounting to about 30%. Lorentz and polarization corrections were applied in the usual manner.

By a procedure related to one used by previous workers,⁹ each recorded number of counts, N , was assigned an estimated standard deviation, $\sigma(N) = [N + (0.01N)^2]^{1/2}$. A total of 2766 reflections for which the net intensity was greater than twice its standard deviation were recorded as observed.

With the final parameters from the refinement with Cu data as a starting point, the refinement was resumed with the Mo data. The first two cycles of least-squares refinement were carried out isotropically and R decreased to 0.11. Next, anisotropic thermal parameters for I, Te, and S atoms were allowed to refine for two cycles ($R = 0.053$) and finally, the thermal parameters for carbon were allowed to vary anisotropically for two cycles, during which R decreased to 0.049.

(9) S. W. Peterson and H. A. Levy, *Acta Crystallogr.*, **10**, 70 (1957).

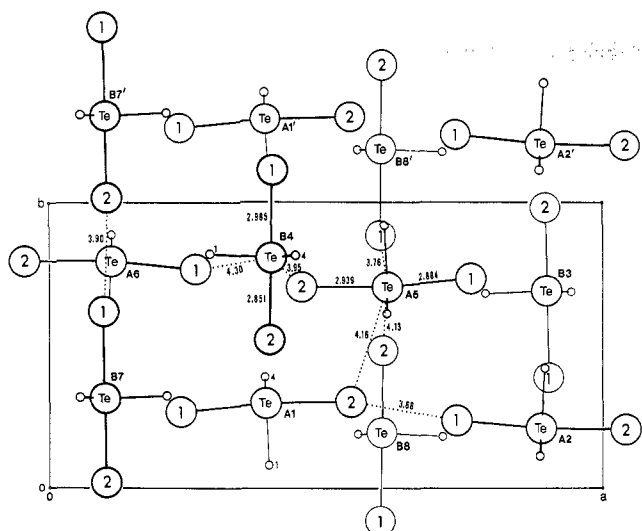


Figure 1.—Projection of a selected portion of the crystal structure of $C_4H_8STeI_2$ down the c axis. The labels A1, B8, etc., designate the crystallographically nonequivalent molecules A and B, and the eight equivalent positions of the space group $Pbca$: (1) x, y, z ; (2) $1/2 + x, 1/2 - y, 1 - z$; (3) $1 - x, 1/2 + y, 1/2 - z$; (4) $1/2 - x, 1 - y, 1/2 + z$; (5) $1 - x, 1 - y, 1 - z$; (6) $1/2 - x, 1/2 + y, z$; (7) $x, 1/2 - y, 1/2 + z$; (8) $1/2 + x, y, 1/2 - z$. For purposes of clarity, the atoms S, C(2), C(3), and all hydrogen atoms are omitted.

At this point a three-dimensional difference Fourier summation was computed, both as a check on the refined structure and for possible location of the hydrogen atoms. The maximum fluctuations on this summation were $\pm 0.8 e \text{ \AA}^{-3}$ or ± 1.6 times the esd of the electron density, $0.5 e \text{ \AA}^{-3}$. Thus, the anticipated diffraction effect of a hydrogen atom was about at the noise level, and while nine of the sixteen hydrogen atoms showed up rather clearly at reasonable positions and with maxima in the range $0.4\text{--}0.6 e \text{ \AA}^{-3}$, information on others was less convincing. A set of hydrogen positions was computed and nine of the locations agreed reasonably well with the maxima mentioned above. Although attempts to refine the hydrogen positions with isotropic B values held constant at 6 \AA^2 were unsuccessful, inclusion of the hydrogen atoms at the calculated parameters given in Table I(b) in a structure factor calculation did reduce R slightly, to 0.048.

A set of final structure factors for all measured reflections was calculated on the basis of the final positional and thermal parameters given in Tables I and II. The resulting values of F_o are listed with the respective F_c values in Table III. The value of R for these 4296 reflections, including 1530 designated as "unobserved," is 0.087.

The atomic scattering curve used for iodine was based on the I^0 values of Doyle and Turner.¹⁰ These authors did not give a curve for tellurium so their values for Sb^0 and I^0 were averaged to give a curve for Te^0 . The curve used for sulfur was that for S^0 given in Table 3.3.1A of the "International Tables for X-Ray Crystallography,"¹¹ that for carbon was based on the dia-

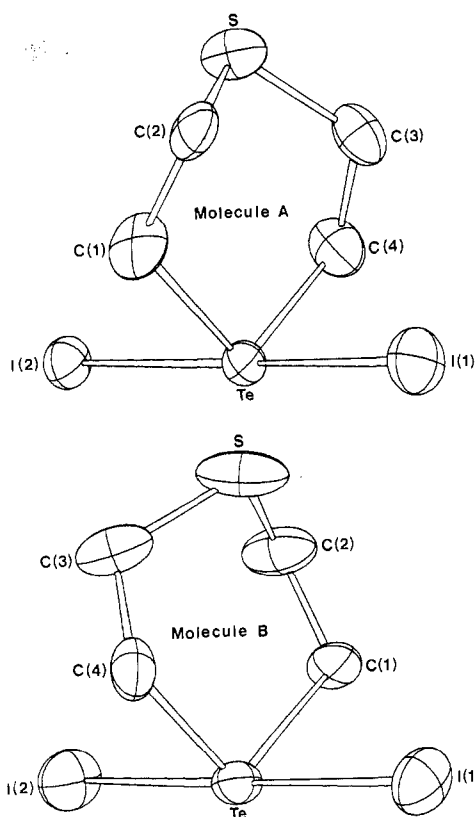


Figure 2.—ORTEP plots of the two independent molecules of $C_4H_8STeI_2$ at angles of about 45° with the respective I-Te-I axes and with the respective C-Te-C planes. The thermal ellipsoids are at the 50% probability level and the hydrogen atoms are omitted.

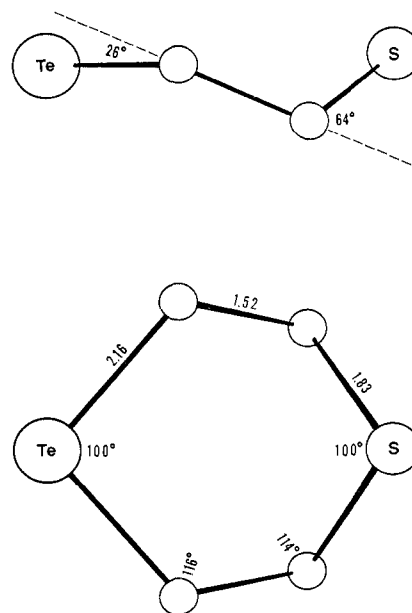


Figure 3.—Views of the structure of the 1,4-thiatellurane ring showing values for the chemically equivalent bond distances (in \AA) and bond angles (in degrees), averaged over the two molecules in the asymmetric unit.

mond values of McWeeny,¹² and the hydrogen curve was based on the bonded values of Stewart, Davidson, and Simpson.¹³ The real part of the anomalous scat-

(10) P. A. Doyle and P. S. Turner, *Acta Crystallogr.*, **A24**, 392 (1968).

(11) "International Tables for X-Ray Crystallography," Vol. III, The Kynoch Press, Birmingham, England, 1962.

(12) R. McWeeny, *Acta Crystallogr.*, **7**, 180 (1954).

(13) R. F. Stewart, E. R. Davidson, and W. T. Simpson, *J. Chem. Phys.*, **42**, 3178 (1965).

TABLE III
OBSERVED AND CALCULATED STRUCTURE FACTORS^a

Table with multiple columns of numerical data representing observed and calculated structure factors for various hkl reflections. The table is organized into several sections, each with a header indicating the reflection type (e.g., h0k0, h0k, hkl) and the corresponding observed and calculated values.

TABLE III (Continued)

Table with multiple columns of numerical data, organized into sections labeled H1, H2, H3, H4, H5, H6, H7, H8, H9, H10, H11, H12, H13, H14, H15, H16, H17, H18, H19, H20, H21, H22, H23, H24, H25, H26, H27, H28, H29, H30, H31, H32, H33, H34, H35, H36, H37, H38, H39, H40, H41, H42, H43, H44, H45, H46, H47, H48, H49, H50, H51, H52, H53, H54, H55, H56, H57, H58, H59, H60, H61, H62, H63, H64, H65, H66, H67, H68, H69, H70, H71, H72, H73, H74, H75, H76, H77, H78, H79, H80, H81, H82, H83, H84, H85, H86, H87, H88, H89, H90, H91, H92, H93, H94, H95, H96, H97, H98, H99, H100. Each section contains a list of numbers, some with sub-labels like H1, H2, etc.

TABLE III (Continued)

H ₂ 3, K ₂ 8	7 71 69	8 20*	4	14 19*	-4	1 19*	9	13 71*	-9	1 21*	-17	4 7*	76	7 20*	12	8 20*	-8	H ₂ 3, K ₂ 11	3 27*	7	
0 4 7*	44	8 19*	9	9 20*	-19	15 27*	-43	2 157	160	14 45*	44	2 4*	-58	5 60*	-58	3 31*	-20	4 23*	35	4 10*	-99
1 17	41	10 23*	4	11 28*	-33	17 20*	16	174	176	15 27*	-8	4 8*	-95	7 3*	32	5 141*	130	11 10*	-104	3 21*	-27
2 19*	24	11 52*	54	12 37	26	18 22*	-14	4 18*	-19	5 24*	16	8 81*	75	4 140*	-141	6 170*	-16	7 170*	16	7 20*	-5
3 18*	-20	12 53*	-3	13 19*	-9	14 17*	-11	5 10*	106	6 21*	-4	9 20*	16	7 170*	16	7 170*	16	7 170*	16	7 170*	16
4 44*	44	13 52*	61	14 21*	-5	1 18*	8	8 119	120	2 1*	-64	10 70*	-49	11 63*	-80	9 47*	-35	2 69*	-63	7 19*	-12
5 43*	-24	14 19*	15	15 21*	-6	2 18*	-21	3 18*	-21	4 1*	-1	12 6*	-2	13 2*	6	13 6*	37	3 158*	31	4 20*	-4
6 19*	38	15 18*	16	16 21*	-11	3 18*	-2	10 142	138	4 37*	-46	10 72*	-43	11 117	119	4 49	50	9 20*	-2	2 41*	-30
7 9*	38	16 24*	21	17 21*	6	4 17*	7	11 37	-44	5 32*	-21	12 6*	-2	13 2*	6	12 101	-108	5 46	28	10 20*	-11
8 10*	19	17 21*	9	18 21*	9	5 17*	-4	12 72	74	6 4*	-15	13 2*	6	14 57	50	12 101	-108	5 46	28	10 20*	-11
9 18*	2	18 21*	9	19 21*	9	6 17*	-4	13 72	74	7 4*	-15	14 57	50	13 2*	6	15 44	-58	8 21*	11	11 20*	-7
10 24*	19	19 21*	9	20 21*	9	7 17*	-4	14 57	50	8 4*	-15	15 44	-58	9 2*	6	16 21*	21	14 44	-29	9 20*	-11
11 18*	-5	20 21*	9	21 21*	9	8 17*	-4	15 44	-58	9 2*	6	17 21*	21	16 21*	21	17 21*	21	15 44	-29	10 20*	-11
12 19*	2	21 21*	9	22 21*	9	9 17*	-4	16 21*	21	10 2*	6	18 21*	21	18 21*	21	19 21*	21	16 21*	21	11 20*	-7
13 19*	20	22 21*	9	23 21*	9	10 17*	-4	17 21*	21	11 2*	6	19 21*	21	20 21*	21	21 21*	21	17 21*	21	12 22*	-11
14 33*	-14	23 21*	9	24 21*	9	11 20*	-25	12 44*	-43	12 44*	-43	13 20*	2	13 20*	2	14 44*	-43	18 21*	21	13 20*	-11
15 19*	-5	24 21*	9	25 21*	9	12 20*	-25	13 20*	2	14 44*	-43	15 20*	2	16 21*	21	15 44*	-29	19 21*	21	14 44*	-29
16 20*	18	25 21*	9	26 21*	9	13 20*	2	14 44*	-43	16 21*	21	17 21*	21	18 21*	21	17 21*	21	20 21*	21	15 44*	-29
17 25*	4	26 21*	9	27 21*	9	14 44*	-29	15 44*	-29	16 21*	21	18 21*	21	19 21*	21	18 21*	21	21 21*	21	16 21*	-11
18 20*	5	27 21*	9	28 21*	9	15 44*	-29	16 21*	21	17 21*	21	19 21*	21	20 21*	21	19 21*	21	22 21*	21	17 21*	-11
19 24*	9	28 21*	9	29 21*	9	16 21*	21	18 21*	21	20 21*	21	21 21*	21	22 21*	21	20 21*	21	23 21*	21	18 21*	-11
20 24*	9	29 21*	9	30 21*	9	17 21*	21	19 21*	21	21 21*	21	22 21*	21	23 21*	21	21 21*	21	24 21*	21	19 21*	-11
21 29*	8	30 21*	9	31 21*	9	18 21*	21	20 21*	21	22 21*	21	23 21*	21	24 21*	21	22 21*	21	25 21*	21	20 21*	-11
22 19*	1	31 21*	9	32 21*	9	19 21*	21	21 21*	21	23 21*	21	24 21*	21	25 21*	21	23 21*	21	26 21*	21	21 21*	-11
23 33*	31	32 21*	9	33 21*	9	20 21*	21	22 21*	21	24 21*	21	25 21*	21	26 21*	21	24 21*	21	27 21*	21	22 21*	-11
24 29*	-23	33 21*	9	34 21*	9	21 21*	21	23 21*	21	25 21*	21	26 21*	21	27 21*	21	25 21*	21	28 21*	21	23 21*	-11
25 40*	-29	34 21*	9	35 21*	9	22 21*	21	24 21*	21	26 21*	21	27 21*	21	28 21*	21	26 21*	21	29 21*	21	24 21*	-11
26 29*	-15	35 21*	9	36 21*	9	23 21*	21	25 21*	21	27 21*	21	28 21*	21	29 21*	21	27 21*	21	30 21*	21	25 21*	-11
27 19*	1	36 21*	9	37 21*	9	24 21*	21	26 21*	21	28 21*	21	29 21*	21	30 21*	21	28 21*	21	31 21*	21	26 21*	-11
28 27*	4	37 21*	9	38 21*	9	25 21*	21	27 21*	21	29 21*	21	30 21*	21	31 21*	21	29 21*	21	32 21*	21	27 21*	-11
29 27*	25	38 21*	9	39 21*	9	26 21*	21	28 21*	21	30 21*	21	31 21*	21	32 21*	21	30 21*	21	33 21*	21	28 21*	-11
30 34*	-28	39 21*	9	40 21*	9	27 21*	21	29 21*	21	31 21*	21	32 21*	21	33 21*	21	31 21*	21	34 21*	21	29 21*	-11
31 11*	18	40 21*	9	41 21*	9	28 21*	21	30 21*	21	32 21*	21	33 21*	21	34 21*	21	32 21*	21	35 21*	21	30 21*	-11
32 22*	-18	41 21*	9	42 21*	9	29 21*	21	31 21*	21	33 21*	21	34 21*	21	35 21*	21	33 21*	21	36 21*	21	31 21*	-11
33 19*	18	42 21*	9	43 21*	9	30 21*	21	32 21*	21	34 21*	21	35 21*	21	36 21*	21	34 21*	21	37 21*	21	32 21*	-11
34 20*	-32	43 21*	9	44 21*	9	31 21*	21	33 21*	21	35 21*	21	36 21*	21	37 21*	21	35 21*	21	38 21*	21	33 21*	-11
35 17*	-10	44 21*	9	45 21*	9	32 21*	21	34 21*	21	36 21*	21	37 21*	21	38 21*	21	36 21*	21	39 21*	21	34 21*	-11
36 10*	7	45 21*	9	46 21*	9	33 21*	21	35 21*	21	37 21*	21	38 21*	21	39 21*	21	37 21*	21	40 21*	21	35 21*	-11
37 10*	7	46 21*	9	47 21*	9	34 21*	21	36 21*	21	38 21*	21	39 21*	21	40 21*	21	38 21*	21	41 21*	21	36 21*	-11
38 19*	49	47 21*	9	48 21*	9	35 21*	21	37 21*	21	39 21*	21	40 21*	21	41 21*	21	39 21*	21	42 21*	21	37 21*	-11
39 19*	21	48 21*	9	49 21*	9	36 21*	21	38 21*	21	40 21*	21	41 21*	21	42 21*	21	40 21*	21	43 21*	21	38 21*	-11
40 34*	-40	49 21*	9	50 21*	9	37 21*	21	39 21*	21	41 21*	21	42 21*	21	43 21*	21	41 21*	21	44 21*	21	39 21*	-11
41 3 19*	49	50 21*	9	51 21*	9	38 21*	21	40 21*	21	42 21*	21	43 21*	21	44 21*	21	42 21*	21	45 21*	21	40 21*	-11
42 37*	-39	51 21*	9	52 21*	9	39 21*	21	41 21*	21	43 21*	21	44 21*	21	45 21*	21	43 21*	21	46 21*	21	41 21*	-11
43 50*	-25	52 21*	9	53 21*	9	40 21*	21	42 21*	21	44 21*	21	45 21*	21	46 21*	21	44 21*	21	47 21*	21	42 21*	-11
44 22*	7	53 21*	9	54 21*	9	41 21*	21	43 21*	21	45 21*	21	46 21*	21	47 21*	21	45 21*	21	48 21*	21	43 21*	-11
45 24*	19	54 21*	9	55 21*	9	42 21*	21	44 21*	21	46 21*	21	47 21*	21	48 21*	21	46 21*	21	49 21*	21	44 21*	-11
46 24*	-19	55 21*	9	56 21*	9	43 21*	21	45 21*	21	47 21*	21	48 21*	21	49 21*	21	47 21*	21	50 21*	21	45 21*	-11
47 19*	26	56 21*	9	57 21*	9	44 21*	21	46 21*	21	48 21*	21	49 21*	21	50 21*	21	48 21*	21	51 21*	21	46 21*	-11
48 10*	-27	57 21*	9	58 21*	9	45 21*	21	47 21*	21	49 21*	21	50 21*	21	51 21*	21	49 21*	21	52 21*	21	47 21*	-11
49 12*	19*	58 21*	9	59 21*	9	46 21*	21	48 21*	21	50 21*	21	51 21*	21	52 21*	21	50 21*	21	53 21*	21	48 21*	-11
50 26*	-34	59 21*	9	60 21*	9	47 21*	21	49 21*	21	51 21*	21	52 21*	21	53 21*	21	51 21*	21	54 21*	21	49 21*	-11
51 12*	19*	60 21*	9	61 21*	9	48 21*	21	50 21*	21	52 21*	21	53 21*	21	54 21*	21	52 21*	21	55 21*	21	50 21*	-11
52 13*	22	61 21*	9	62 21*	9	49 21*	21	51 21*	21	53 21*	21	54 21*	21	55 21*	21	53 21*	21	56 21*	21	51 21*	-11
53 19*	-14	62 21*	9	63 21*	9	50 21*	21	52 21*	21	54 21*	21	55 21*	21	56 21*	21	54 21*	21	57 21*	21	52 21*	-11
54 15*	-24	63 21*	9	64 21*	9	51 21*	21	53 21*	21	55 21*	21	56 21*	21	57 21*	21	55 21*	21	58 21*	21	53 21*	-11
55 16*	32	64 21*	9	65 21*	9	52 21*	21	54 21*	21	56 21*	21	57 21*	21	58 21*	21	56 21*	21	59 21*	21	54 21*	-11
56 40*	-32	65 21*	9	66 21*	9	53 21*	21	55 21*	21												

carbon atoms are coplanar within the standard deviations of the positional parameters. Individual atom displacements from the least-squares planes through the carbon atoms are given in Table VII. The planes of the C(1)-Te-C(4) systems make dihedral angles of $27(1)^\circ$ in molecule A and $25(1)^\circ$ in molecule B with the respective least-squares planes through the carbon atoms. In contrast, the dihedral angles that these least-squares planes make with the C(2)-S-C(3) systems are $64(2)^\circ$ in molecule A and $63(2)^\circ$ in molecule B. This geometry has the effect of bringing the sulfur atoms close to the planes of the C-Te-C systems. These distances are $0.063(7) \text{ \AA}$ in molecule A and $0.122(7) \text{ \AA}$ in molecule B. In each case, the S atom is on the opposite side of the C-Te-C plane from the locations of C(2) and C(3).

The molecular structure of 1,4-thiatellurane diiodide is very much like that of 1,4-thiaselenane dibromide,¹⁴ except for displacement of the iodine atoms caused by intermolecular interactions. The molecules approximate mirror symmetry through the I, Te, and S atoms and the configuration about tellurium approximates a trigonal bipyramid with C(1), C(4), and an unshared pair of electrons in equatorial positions and the iodine atoms in axial positions. This type of geometry about tellurium has been observed previously in diphenyltellurium dibromide,¹⁵ dimethyltellurium dichloride,¹⁶ and *p,p*-dichlorodiphenyltellurium diiodide,⁵ as well as in numerous selenium compounds of the general type R_2SeX_2 .¹⁴

The main features of the intermolecular bonding are shown in Figure 1. Two sets of loosely bonded linear polymers are formed and there are significant "cross links" between the sets. In each set, systems of the type $\cdots I(1)-Te-I(2)\cdots I(1)-Te-I(2)\cdots$ are found. In the A molecules these systems are lined up along the crystallographic *a* axis while in the B molecules they are lined up along *b*. The intermolecular I(1) \cdots I(2) separations in the chains are 3.66 \AA in the A set and 3.90 \AA in the B set whereas the van der Waals radius of 2.15 \AA ¹⁷ for iodine would suggest a normal, nonbonded contact of 4.30 \AA . The cross linking between chains consists of secondary bonding of the tellurium atoms in one chain with iodine atoms in another. These secondary bonds increase the equatorial ligancy about tellurium to five in molecule A and to four in molecule B. Thus, the configuration about each tellurium atom approximates a pentagonal bipyramid, but for Te(B), one equatorial iodine atom is absent. The bonding arrangements about the tellurium atoms are shown in Figure 4. Each iodine atom forms a strong bond with the tellurium atom in its own molecule and a secondary bond with an iodine atom in a neighboring molecule in the same chain, to give a nearly linear group. In each

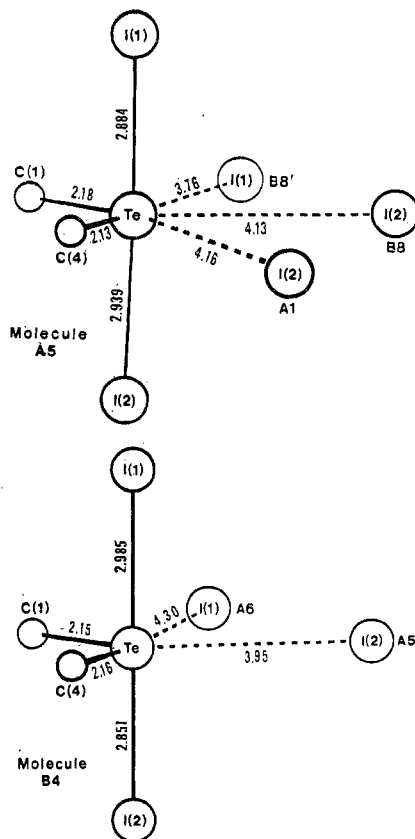
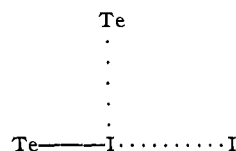


Figure 4.—Views showing the bonding configuration about Te(A5) and Te(B4). Primary (intramolecular) bonds are shown by full lines, secondary (intermolecular) bonds by dotted lines. The bond angles about these tellurium atoms are given in Table VIII and the atom numbering is the same as used in that table and in Figure 1.

case, the central iodine atom makes a secondary bond with a tellurium atom of the opposite set, *i.e.*, I(A) with Te(B) and I(B) with Te(A), so that the Te-I \cdots Te bond angle is in the range 103 to 125° . The result is a nearly planar group of two iodine and two tellurium atoms arranged thus



In one case, I(2) of set A, an additional tellurium atom forms a secondary bond with the central iodine atom so that the second Te-I \cdots Te angle is 108° and the Te \cdots I \cdots Te angle is 97° . Further details of the coordination about the tellurium and iodine atoms are given in Table VIII.

An examination of the data in Table VIII shows that there is excellent correlation between the strength of the bond (as indicated by bond distance) that an iodine atom forms with the Te atom in its own molecule and the strengths of the secondary bonds which it forms. Thus, I(1) of molecule B forms the longest primary bond with Te, 2.985 \AA , but it also forms the shortest intermolecular bond with Te at 3.76 \AA . On the other hand,

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(16) G. D. Christofferson, R. A. Sparks, and J. D. McCullough, *ibid.*, **11**, 782 (1958).

(17) L. Pauling, "Nature of the Chemical Bond," 3rd ed, Cornell University Press, Ithaca, N. Y., 1960, p 260.

TABLE VIII
 COORDINATION ABOUT TELLURIUM AND IODINE ATOMS IN $C_4H_8STeI_2$

Angle, deg ^b	Bonded atom	Dist, Å	Angle, deg ^b	Bonded atom	Dist, Å
	About Te(A5) ^a			About I(1) (A6)	
174.9	I(1) (A5) ax ^c	2.884	174.9	Te (A6) ax	2.884
87.4 ^d	I(2) (A5) ax	2.939	58.9	I(2) (A5) ax	3.66
100	C(1) (A5) eq	2.18	124.9	Te (B4)	4.30
61.9	C(4) (A5) eq	2.13		Te (A6) ax	2.884
63.0	I(2) (A1) eq	4.16		About I(2) (A5) ^e	
59.0	I(2) (B8) eq	4.13	167.9	Te (A5) ax	2.939
78.1	I(1) (B8') eq	3.76	68.6	I(1) (A6) ax	3.66
	C(1) (A5) ^e eq	2.18	102.9	Te (B4)	3.95
	About Te(B4)			Te (A5) ax	2.939
178.1	I(1) (B4) ax	2.985		About I(1) (B8')	
91.6 ^f	I(2) (B4) ax	2.851	175.3	Te (B8') ax	2.985
100	C(1) (B4) eq	2.15	65.3	I(2) (B8) ax	3.90
138.9	C(4) (B4) eq	2.16	118.3	Te (A5)	3.76
52.5	I(2) (A5) eq	3.95		Te (B8') ax	2.985
66.5	I(1) (A6) eq	4.30	174.2	About I(2) (B8)	
	C(1) (B4) eq	2.15	55.7	Te (B8) ax	2.851
			118.9	I(1) (B8') ax	3.90
				Te (A5)	4.13
				Te (B8) ax	2.851

^a The atom and molecule numbering is keyed to Figure 1. ^b The angles are those formed at the central atom by the bonds to the two bracketed atoms. ^c The atoms are designated ax for axial and eq for equatorial where this designation is useful. ^d The ten axial-equatorial angles at Te(A) range from 72.3 to 112.8°. ^e One atom is repeated in each coordination group to facilitate the designation of angles. ^f The eight axial-equatorial angles at Te(B) range from 77.3 to 104.1°. ^g An additional tellurium atom, Te(A1), at a distance of 4.16 Å from the central atom, is located as indicated in the text.

I(2) of molecule B forms the shortest primary bond with Te, 2.851 Å, and its only intermolecular bond with Te has a length of 4.13 Å.

The thermal ellipsoids of the iodine atoms show an interesting correlation with the observed structure. In molecule A, the I-Te-I axis is lined up with the crystallographic *a* axis and the major vibrations of the iodine atoms are perpendicular to *a* and hence to the direction of the Te-I bonds. In molecule B, where the alignment is along *b*, the major vibrations of the iodine atoms are perpendicular to *b* and (again) to the direc-

tion of the Te-I bonds. This effect is smaller in I(2)A than in the other three iodine atoms, and inspection of the structure shows that this atom is more restricted in its vibrations by the larger number of secondary bonds and by their directions.

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